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Hexakis(μ_2-2-aminoethanethiolato)-triiron(III) tris(perchlorate)

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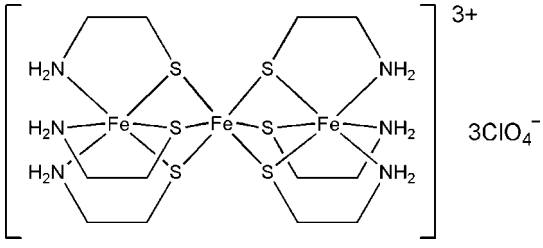
Key indicators: single-crystal X-ray study; T = 200 K; mean σ(C–C) = 0.003 Å; disorder in main residue; R factor = 0.032; wR factor = 0.100; data-to-parameter ratio = 14.3.

In the title salt, $[Fe_3(C_2H_6NS)_6](ClO_4)_3$, the trinuclear cation lies on a special position of $\overline{3}$ site symmetry; the central Fe atom is coordinated by six thiolate groups from the two flanking *fac*-(*S*)-$[Fe(C_2H_6NS)_3]$ units. In the flanking units, the three C_2H_6NS groups each chelate to the metal atom. The cations interact with the perchlorate anions through weak N—H···O hydrogen bonds resulting in a three-dimensional network. In the asymmetric unit two cations are present, one of which is disordered over two positions with occupancies of 0.75 and 0.25.

Related literature

For related structures, see: Busch & Jicha (1962); Heeg *et al.* (1985); Mahboob *et al.* (2004); Marsh *et al.* (1986); Matsuura *et al.* (2006).



Experimental

Crystal data

$[Fe_3(C_2H_6NS)_6](ClO_4)_3$
M_r = 922.73
Trigonal, $R\overline{3}$
a = 14.2852 (6) Å
c = 26.2187 (8) Å
V = 4633.6 (2) Å^3
Z = 6
Mo $K\alpha$ radiation
μ = 2.12 mm^{-1}
T = 200 (2) K
0.20 × 0.20 × 0.10 mm

Data collection

Rigaku R-AXIS RAPID diffractometer
Absorption correction: multi-scan (*ABSCOR*; Higashi, 1995)
T_{min} = 0.677, T_{max} = 0.816
15327 measured reflections
2365 independent reflections
2144 reflections with $I > 2\sigma(I)$
R_{int} = 0.053

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.032
$wR(F^2)$ = 0.100
S = 1.39
2365 reflections
165 parameters
6 restraints
H-atom parameters constrained
$\Delta\rho_{max}$ = 1.09 e Å^{-3}
$\Delta\rho_{min}$ = −0.31 e Å^{-3}

Table 1
Selected bond lengths (Å).

Fe1—S1	2.2764 (5)	Fe4—N2	2.026 (4)
Fe2—N1	2.0482 (18)	Fe4—N2*B*	2.059 (11)
Fe2—S1	2.2434 (6)	Fe4—S2*B*	2.229 (2)
Fe3—S2*B*	2.281 (2)	Fe4—S2	2.2535 (8)
Fe3—S2	2.2869 (7)		

Table 2
Hydrogen-bond geometry (Å, °).

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N1—H1···O2	0.92	2.28	3.130 (3)	154
N1—H2···O3^{iii}	0.92	2.40	3.162 (3)	140
N2—H7···O1^{ii}	0.92	2.30	3.110 (4)	147
N2—H8···O2	0.92	2.39	3.274 (4)	161
N2*B*—H13···O2	0.92	2.41	2.984 (12)	121
N2*B*—H14···O1	0.92	2.27	3.112 (11)	152

Symmetry codes: (ii) $-x+y, -x+1, z$; (iii) $x-y+\frac{1}{3}, x+\frac{2}{3}, -z+\frac{2}{3}$.

Data collection:*PROCESS-AUTO* (Rigaku, 1998); cell refinement: *PROCESS-AUTO*; data reduction: *CrystalStructure* (Rigaku/MSC, 2004); program(s) used to solve structure: *SIR97* (Altomare *et al.*, 1999); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *Yadokari-XG* (Wakita, 2000).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NG2504).

References

Altomare, A., Burla, M. C., Camalli, M., Cascarano, G. L., Giacovazzo, C., Guagliardi, A., Moliterni, A. G. G., Polidori, G. & Spagna, R. (1999). *J. Appl. Cryst.* **32**, 115–119.
Busch, D. C. & Jicha, D. C. (1962). *Inorg. Chem.* **1**, 884–887.
Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
Heeg, M. J., Blinn, E. L. & Deutsch, E. (1985). *Inorg. Chem.* **24**, 1118–1120.
Higashi, T. (1995). *ABSCOR*. Rigaku Corporation, Tokyo, Japan.
Mahboob, N., Miyashita, Y., Yamada, Y., Fujisawa, K. & Okamoto, K. (2004). *Inorg. Chim. Acta*, **357**, 75–82.
Marsh, R. E., Heeg, M. J. & Deutsch, E. (1986). *Inorg. Chem.* **25**, 118.
Matsuura, N., Igashira-Kamiyama, A., Kawamoto, T. & Konno, T. (2006). *Inorg. Chem.* **45**, 401–408.

Rigaku (1998). *PROCESS-AUTO*. Rigaku Corporation, Tokyo, Japan.
Rigaku/MSC (2004). *CrystalStructure*. Rigaku/MSC, The Woodlands, Texas, USA.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Wakita, K. (2000). *Yadokari-XG*. Department of Chemistry, Graduate School of Science, The University of Tokyo, Japan.

supplementary materials

Acta Cryst. (2008). E**64**, m1519-m1520 [**doi:10.1107/S1600536808036167**]

Hexakis(μ_2-2-aminoethanethiolato)triiron(III) tris(perchlorate)

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Comment

Due to the high Lewis basicity of thiolate groups, a great number of thiolate-bridged complexes have been synthesized up to now. 2-Aminoethanethiolate (aet) is the simplest N,*S*-chelating ligand that has been used for the formation of S-bridged polynuclear structures. For example, it has been shown that the aet ligand reacts with the octahedral metal ions to give linear-type S-bridged trinuclear structures, such as Co^{III}_3 (Busch and Jicha, 1962; Heeg *et al.*, 1985; Marsh *et al.*, 1986), Rh^{III}_3 (Mahboob *et al.*, 2004), and Ru^{III}_3 (Matsuura *et al.*, 2006). In this paper, we report on the crystal structure of the title compound (I), which was obtained by the reaction of aet and $Fe(ClO_4)_3$. The asymmetric unit of the compound (I) contains two complex cations having a threefold rotation-inversion axis and one perchlorate anion. One of the complex cations is disordered over two positions with occupancies of 0.75 and 0.25. In the complex cation of (I), two *fac*(*S*)-[$Fe(aet)_3$] units coordinate to a central Fe atom through thiolato bridges to form a linear-type trinuclear structure. Each terminal Fe atom is in an N_3S_3 octahedral environment, whereas the central Fe atom is in an S_6 octahedral environment. Considering the charge balance, it is assumed that all Fe atoms have a +III oxidation state.

Experimental

To a solution containing 2-aminoethanethiol hydrochloride (0.11 g, 1 mmol) in 20 ml of methanol/CH_2Cl_2 (1:1) was added a solution of Et_3N (0.10 g, 1 mmol) in 10 ml of methanol and a solution of $Fe(ClO_4)_3.6H_2O$ (0.09 g, 0.2 mmol) in 2 ml of methanol. The resulting dark brown solution was stood at room temperature overnight to give black crystals, which was filtered and washed with methanol. Yield: 31 mg (50% based on Fe). Anal. Calcd for $[Fe_3(aet)_6](ClO_4)_3 = C_{12}H_{36}Cl_3Fe_3N_6O_{12}S_6$: C, 15.62; H, 3.93; N, 9.11%. Found: C, 15.82; H, 3.88; N, 9.00%.

Refinement

H atoms bonded to C and N atoms were placed at calculated positions [C—H = 0.99 and N—H = 0.92 Å] and refined as riding with $U_{iso}(H) = 1.2U_{eq}$ (C,*N*). One cationic part was disordered over two positions (S2, N2, C3, C4 and S2B, N2B, C3B, C4B) and refined with site occupancies of 0.75 and 0.25. The C3B atom in a minor component was restrained based on ISOR.

Figures

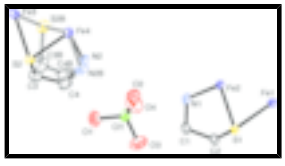


Fig. 1. The asymmetric unit of (I), with atom labels and 50% probability displacement ellipsoids for non-H atoms. Hydrogen atoms are omitted for clarity.

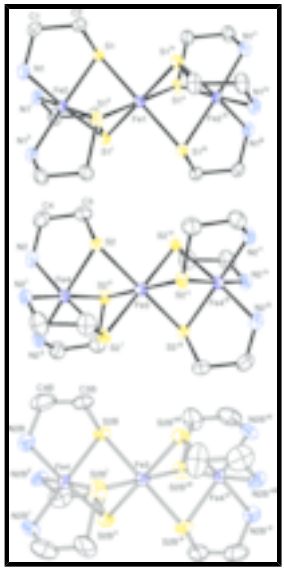


Fig. 2. The cation units of (I). The molecule with open bonds is the minor component of the disordered unit. Symmetry codes: (i) -*y* + 1, *x*-*y* + 1, *z*; (ii) -*x* + *y*, -*x* + 1, *z*; (iii) -*x* + 2/3, -*y* + 4/3, -*z* + 1/3; (iv) *y* - 1/3, -*x* + *y*+1/3, -*z* + 1/3; (v) *x*-*y* + 2/3, *x* + 1/3, -*z* + 1/3; (vi) -*x* + 2/3, -*y* + 4/3, -*z* + 4/3; (vii) *y* - 1/3, -*x* + *y*+1/3, -*z* + 4/3; (viii) *x*-*y* + 2/3, *x* + 1/3, -*z* + 4/3.

Hexakis(μ$_2$-2-aminoethanethiolato)triiron(III) tris(perchlorate)

Crystal data

$[Fe_3(C_2H_6NS)_6](ClO_4)_3$	$Z = 6$
$M_r = 922.73$	$F_{000} = 2826$
Trigonal, $R\bar{3}$	$D_x = 1.984$ Mg m^{-3}
Hall symbol: -R 3	Mo $K\alpha$ radiation λ = 0.71075 Å
$a = 14.2852$ (6) Å	Cell parameters from 12827 reflections
$b = 14.2852$ Å	θ = 3.4–27.4º
$c = 26.2187$ (8) Å	μ = 2.12 mm^{-1}
α = 90º	$T = 200$ (2) K
β = 90º	Prism, black
γ = 120º	0.20 × 0.20 × 0.10 mm
$V = 4633.6$ (2) Å^3	

Data collection

Rigaku R-AXIS RAPID diffractometer	2365 independent reflections
Radiation source: fine-focus sealed tube	2144 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{int} = 0.053$
Detector resolution: 10.00 pixels mm^{-1}	$\theta_{max} = 27.4$º
$T = 200$(2) K	$\theta_{min} = 3.4$º
ω scans	$h = -18 \rightarrow 16$
Absorption correction: multi-scan (ABSCOR; Higashi, 1995)	$k = -18 \rightarrow 17$
$T_{min} = 0.677$, $T_{max} = 0.816$	$l = -33 \rightarrow 33$
15327 measured reflections	

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.032$	H-atom parameters constrained
$wR(F^2) = 0.100$	$w = 1/[\sigma^2(F_o^2) + (0.0456P)^2 + 3.0528P]$ where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.39$	$(\Delta/\sigma)_{max} = 0.001$
2365 reflections	$\Delta\rho_{max} = 1.09$ e Å^{-3}
165 parameters	$\Delta\rho_{min} = -0.31$ e Å^{-3}
6 restraints	Extinction correction: none
Primary atom site location: structure-invariant direct methods	

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger. The 2-aminoethanethiolate ligand of one unit containing Fe3 and Fe4 atoms is disordered over two positions with the occupancies of 0.75 and 0.25. The C3B atom in the minor component is restrained based on ISOR.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}	Occ. (<1)
Fe1	0.3333	0.6667	0.1667	0.01712 (18)	
Fe2	0.3333	0.6667	0.269870 (17)	0.01673 (15)	
S1	0.21510 (4)	0.53292 (4)	0.219319 (19)	0.02149 (15)	
N1	0.20358 (15)	0.64420 (15)	0.31253 (7)	0.0255 (4)	
H1	0.2151	0.6336	0.3460	0.031*	
H2	0.1986	0.7059	0.3110	0.031*	
C1	0.09906 (18)	0.5506 (2)	0.29524 (9)	0.0308 (5)	
H3	0.0382	0.5595	0.3071	0.037*	
H4	0.0903	0.4828	0.3100	0.037*	
C2	0.09795 (17)	0.5441 (2)	0.23732 (9)	0.0302 (5)	
H5	0.1026	0.6098	0.2222	0.036*	
H6	0.0306	0.4802	0.2253	0.036*	
Fe3	0.3333	0.6667	0.6667	0.02171 (19)	
Fe4	0.3333	0.6667	0.563330 (19)	0.02155 (16)	
S2	0.20595 (6)	0.53902 (6)	0.61393 (3)	0.02341 (18)	0.75

N2	0.2103 (3)	0.6563 (3)	0.52058 (14)	0.0300 (8)	0.75
H7	0.2163	0.7234	0.5193	0.036*	0.75
H8	0.2176	0.6377	0.4878	0.036*	0.75
C3	0.0975 (3)	0.5650 (4)	0.59679 (17)	0.0394 (9)	0.75
H9	0.0267	0.5041	0.6076	0.047*	0.75
H10	0.1083	0.6319	0.6135	0.047*	0.75
C4	0.1011 (4)	0.5773 (4)	0.5393 (2)	0.0425 (11)	0.75
H11	0.0776	0.5062	0.5233	0.051*	0.75
H12	0.0496	0.6011	0.5289	0.051*	0.75
S2B	0.2622 (3)	0.7412 (2)	0.61334 (10)	0.0335 (5)	0.25
N2B	0.1951 (9)	0.6067 (9)	0.5199 (5)	0.035 (3)	0.25
H13	0.2127	0.6372	0.4879	0.041*	0.25
H14	0.1655	0.5332	0.5163	0.041*	0.25
C3B	0.1180 (11)	0.6301 (17)	0.5950 (6)	0.065 (4)	0.25
H15	0.1036	0.5589	0.6078	0.078*	0.25
H16	0.0638	0.6462	0.6098	0.078*	0.25
C4B	0.1130 (14)	0.6295 (15)	0.5432 (8)	0.060 (5)	0.25
H17	0.1257	0.7006	0.5311	0.072*	0.25
H18	0.0399	0.5737	0.5321	0.072*	0.25
Cl1	0.17461 (5)	0.41937 (4)	0.42062 (2)	0.03377 (17)	
O1	0.12758 (17)	0.38390 (16)	0.47031 (7)	0.0460 (5)	
O2	0.20389 (18)	0.53133 (14)	0.41480 (8)	0.0492 (5)	
O3	0.0981 (2)	0.35574 (17)	0.38241 (8)	0.0604 (6)	
O4	0.26978 (18)	0.41114 (19)	0.41632 (9)	0.0583 (6)	

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Fe1	0.0171 (2)	0.0171 (2)	0.0173 (4)	0.00853 (12)	0.000	0.000
Fe2	0.01850 (19)	0.01850 (19)	0.0132 (3)	0.00925 (9)	0.000	0.000
S1	0.0233 (3)	0.0207 (3)	0.0166 (3)	0.0081 (2)	−0.00003 (18)	0.00068 (17)
N1	0.0293 (10)	0.0302 (10)	0.0174 (9)	0.0151 (8)	0.0022 (7)	0.0004 (7)
C1	0.0224 (11)	0.0387 (13)	0.0235 (12)	0.0094 (10)	0.0051 (8)	0.0034 (9)
C2	0.0180 (10)	0.0410 (13)	0.0235 (11)	0.0088 (10)	0.0020 (8)	0.0009 (9)
Fe3	0.0190 (2)	0.0190 (2)	0.0272 (4)	0.00950 (12)	0.000	0.000
Fe4	0.0229 (2)	0.0229 (2)	0.0189 (3)	0.01145 (10)	0.000	0.000
S2	0.0220 (4)	0.0236 (4)	0.0205 (4)	0.0082 (3)	−0.0005 (3)	−0.0008 (3)
N2	0.0387 (19)	0.033 (2)	0.0200 (15)	0.0197 (19)	−0.0061 (12)	−0.0045 (16)
C3	0.0233 (18)	0.052 (2)	0.040 (2)	0.0165 (19)	−0.0008 (14)	0.001 (2)
C4	0.031 (2)	0.060 (3)	0.036 (2)	0.023 (3)	−0.0093 (15)	−0.007 (2)
S2B	0.0387 (15)	0.0385 (14)	0.0369 (14)	0.0294 (14)	−0.0028 (11)	0.0001 (11)
N2B	0.034 (5)	0.030 (6)	0.037 (5)	0.014 (6)	−0.011 (4)	−0.014 (5)
C3B	0.026 (6)	0.124 (13)	0.053 (8)	0.043 (9)	−0.010 (5)	−0.014 (10)
C4B	0.033 (7)	0.066 (12)	0.082 (12)	0.026 (9)	−0.018 (7)	−0.014 (11)
Cl1	0.0459 (4)	0.0229 (3)	0.0229 (3)	0.0100 (2)	−0.0001 (2)	0.00124 (19)
O1	0.0606 (13)	0.0428 (11)	0.0322 (10)	0.0239 (10)	0.0109 (9)	0.0091 (8)
O2	0.0767 (15)	0.0243 (9)	0.0382 (12)	0.0188 (9)	−0.0039 (10)	0.0027 (7)
O3	0.0780 (16)	0.0390 (11)	0.0502 (13)	0.0188 (11)	−0.0270 (11)	−0.0113 (9)

O4	0.0501 (12)	0.0571 (13)	0.0655 (15)	0.0252 (11)	0.0103 (10)	0.0076 (11)

Geometric parameters (Å, °)

Fe1—S1^{i}	2.2763 (5)	Fe4—N2^{iii}	2.026 (4)
Fe1—S1^{ii}	2.2763 (5)	Fe4—N2^{iv}	2.026 (4)
Fe1—S1^{iii}	2.2764 (5)	Fe4—N2B	2.059 (11)
Fe1—S1	2.2764 (5)	Fe4—N2B^{iii}	2.059 (11)
Fe1—S1^{iv}	2.2764 (5)	Fe4—N2B^{iv}	2.059 (11)
Fe1—S1^{v}	2.2764 (5)	Fe4—S2B	2.229 (2)
Fe2—N1	2.0482 (18)	Fe4—S2B^{iii}	2.229 (2)
Fe2—N1^{iii}	2.0482 (18)	Fe4—S2B^{iv}	2.229 (2)
Fe2—N1^{iv}	2.0482 (18)	Fe4—S2^{iii}	2.2535 (7)
Fe2—S1^{iii}	2.2434 (6)	Fe4—S2	2.2535 (8)
Fe2—S1	2.2434 (6)	Fe4—S2^{iv}	2.2535 (7)
Fe2—S1^{iv}	2.2434 (6)	S2—C3	1.821 (4)
S1—C2	1.821 (2)	N2—C4	1.480 (7)
N1—C1	1.492 (3)	N2—H7	0.9200
N1—H1	0.9200	N2—H8	0.9200
N1—H2	0.9200	C3—C4	1.514 (7)
C1—C2	1.521 (3)	C3—H9	0.9900
C1—H3	0.9900	C3—H10	0.9900
C1—H4	0.9900	C4—H11	0.9900
C2—H5	0.9900	C4—H12	0.9900
C2—H6	0.9900	S2B—C3B	1.929 (16)
Fe3—S2B^{vi}	2.281 (2)	N2B—C4B	1.49 (2)
Fe3—S2B	2.281 (2)	N2B—H13	0.9200
Fe3—S2B^{iv}	2.281 (2)	N2B—H14	0.9200
Fe3—S2B^{iii}	2.281 (2)	C3B—C4B	1.36 (3)
Fe3—S2B^{vii}	2.281 (2)	C3B—H15	0.9900
Fe3—S2B^{viii}	2.281 (2)	C3B—H16	0.9900
Fe3—S2^{viii}	2.2869 (7)	C4B—H17	0.9900
Fe3—S2^{iii}	2.2869 (7)	C4B—H18	0.9900
Fe3—S2	2.2869 (7)	Cl1—O3	1.425 (2)
Fe3—S2^{iv}	2.2869 (7)	Cl1—O4	1.426 (2)
Fe3—S2^{vi}	2.2870 (7)	Cl1—O1	1.4369 (18)
Fe3—S2^{vii}	2.2870 (7)	Cl1—O2	1.4447 (18)
Fe4—N2	2.026 (4)		
S1^{i}—Fe1—S1^{ii}	87.042 (18)	N2—Fe4—N2^{iii}	92.34 (15)
S1^{i}—Fe1—S1^{iii}	180.00 (3)	N2—Fe4—N2^{iv}	92.34 (15)
S1^{ii}—Fe1—S1^{iii}	92.959 (18)	N2^{iii}—Fe4—N2^{iv}	92.34 (15)
S1^{i}—Fe1—S1	92.960 (18)	N2B—Fe4—N2B^{iii}	92.4 (5)
S1^{ii}—Fe1—S1	180.0	N2—Fe4—N2B^{iv}	103.6 (3)

S1^{iii}—Fe1—S1	87.039 (18)	N2^{iii}—Fe4—N2B^{iv}	78.4 (3)
S1^{i}—Fe1—S1^{iv}	92.961 (18)	N2^{iv}—Fe4—N2B^{iv}	17.7 (3)
S1^{ii}—Fe1—S1^{iv}	92.961 (18)	N2B—Fe4—N2B^{iv}	92.4 (5)
S1^{iii}—Fe1—S1^{iv}	87.039 (18)	N2B^{iii}—Fe4—N2B^{iv}	92.4 (5)
S1—Fe1—S1^{iv}	87.039 (18)	N2B—Fe4—S2B	87.3 (3)
S1^{i}—Fe1—S1^{v}	87.041 (18)	N2B^{iii}—Fe4—S2B	91.5 (4)
S1^{ii}—Fe1—S1^{v}	87.040 (18)	N2B^{iv}—Fe4—S2B	176.1 (3)
S1^{iii}—Fe1—S1^{v}	92.959 (18)	N2B—Fe4—S2B^{iii}	176.1 (3)
S1—Fe1—S1^{v}	92.960 (18)	N2B^{iii}—Fe4—S2B^{iii}	87.3 (3)
S1^{iv}—Fe1—S1^{v}	180.0	N2B^{iv}—Fe4—S2B^{iii}	91.5 (4)
N1—Fe2—N1^{iii}	93.02 (7)	S2B—Fe4—S2B^{iii}	88.91 (9)
N1—Fe2—N1^{iv}	93.02 (7)	N2B—Fe4—S2B^{iv}	91.5 (4)
N1^{iii}—Fe2—N1^{iv}	93.02 (7)	N2B^{iii}—Fe4—S2B^{iv}	176.1 (3)
N1—Fe2—S1^{iii}	91.05 (5)	N2B^{iv}—Fe4—S2B^{iv}	87.3 (3)
N1^{iii}—Fe2—S1^{iii}	87.26 (5)	S2B—Fe4—S2B^{iv}	88.91 (9)
N1^{iv}—Fe2—S1^{iii}	175.90 (5)	S2B^{iii}—Fe4—S2B^{iv}	88.91 (9)
N1—Fe2—S1	87.26 (5)	N2—Fe4—S2^{iii}	91.91 (11)
N1^{iii}—Fe2—S1	175.90 (5)	N2^{iii}—Fe4—S2^{iii}	86.94 (11)
N1^{iv}—Fe2—S1	91.05 (5)	N2^{iv}—Fe4—S2^{iii}	175.71 (11)
S1^{iii}—Fe2—S1	88.65 (2)	N2—Fe4—S2	86.94 (11)
N1—Fe2—S1^{iv}	175.90 (5)	N2^{iii}—Fe4—S2	175.71 (11)
N1^{iii}—Fe2—S1^{iv}	91.05 (5)	N2^{iv}—Fe4—S2	91.91 (11)
N1^{iv}—Fe2—S1^{iv}	87.26 (5)	S2^{iii}—Fe4—S2	88.86 (3)
S1^{iii}—Fe2—S1^{iv}	88.65 (2)	N2—Fe4—S2^{iv}	175.71 (11)
S1—Fe2—S1^{iv}	88.65 (2)	N2^{iii}—Fe4—S2^{iv}	91.91 (11)
C2—S1—Fe2	96.07 (8)	N2^{iv}—Fe4—S2^{iv}	86.94 (11)
C2—S1—Fe1	114.36 (8)	S2^{iii}—Fe4—S2^{iv}	88.86 (3)
Fe2—S1—Fe1	73.546 (18)	S2—Fe4—S2^{iv}	88.86 (3)
C1—N1—Fe2	113.33 (14)	C3—S2—Fe4	96.60 (14)
C1—N1—H1	108.9	C3—S2—Fe3	113.93 (15)
Fe2—N1—H1	108.9	Fe4—S2—Fe3	73.27 (2)
C1—N1—H2	108.9	C4—N2—Fe4	114.7 (3)
Fe2—N1—H2	108.9	C4—N2—H7	108.6
H1—N1—H2	107.7	Fe4—N2—H7	108.6
N1—C1—C2	109.46 (17)	C4—N2—H8	108.6
N1—C1—H3	109.8	Fe4—N2—H8	108.6
C2—C1—H3	109.8	H7—N2—H8	107.6
N1—C1—H4	109.8	C4—C3—S2	106.6 (3)
C2—C1—H4	109.8	C4—C3—H9	110.4
H3—C1—H4	108.2	S2—C3—H9	110.4
C1—C2—S1	106.38 (15)	C4—C3—H10	110.4
C1—C2—H5	110.5	S2—C3—H10	110.4
S1—C2—H5	110.5	H9—C3—H10	108.6

C1—C2—H6	110.5	N2—C4—C3	112.4 (4)
S1—C2—H6	110.5	N2—C4—H11	109.1
H5—C2—H6	108.6	C3—C4—H11	109.1
S2B^{vi}—Fe3—S2B	179.998 (1)	N2—C4—H12	109.1
S2B^{vi}—Fe3—S2B^{iv}	93.64 (9)	C3—C4—H12	109.1
S2B—Fe3—S2B^{iv}	86.37 (9)	H11—C4—H12	107.9
S2B^{vi}—Fe3—S2B^{iii}	93.64 (9)	C3B—S2B—Fe4	90.9 (5)
S2B—Fe3—S2B^{iii}	86.37 (9)	C3B—S2B—Fe3	108.1 (6)
S2B^{iv}—Fe3—S2B^{iii}	86.37 (9)	Fe4—S2B—Fe3	73.83 (7)
S2B^{vi}—Fe3—S2B^{vii}	86.36 (9)	C4B—N2B—Fe4	112.1 (9)
S2B—Fe3—S2B^{vii}	93.63 (9)	C4B—N2B—H13	109.2
S2B^{iv}—Fe3—S2B^{vii}	180.00 (10)	Fe4—N2B—H13	109.2
S2B^{iii}—Fe3—S2B^{vii}	93.64 (9)	C4B—N2B—H14	109.2
S2B^{vi}—Fe3—S2B^{viii}	86.36 (9)	Fe4—N2B—H14	109.2
S2B—Fe3—S2B^{viii}	93.63 (9)	H13—N2B—H14	107.9
S2B^{iv}—Fe3—S2B^{viii}	93.64 (9)	C4B—C3B—S2B	106.6 (13)
S2B^{iii}—Fe3—S2B^{viii}	179.998 (1)	C4B—C3B—H15	110.4
S2B^{vii}—Fe3—S2B^{viii}	86.36 (9)	S2B—C3B—H15	110.4
S2^{viii}—Fe3—S2^{iii}	180.00 (5)	C4B—C3B—H16	110.4
S2^{viii}—Fe3—S2	92.77 (3)	S2B—C3B—H16	110.4
S2^{iii}—Fe3—S2	87.23 (3)	H15—C3B—H16	108.6
S2^{viii}—Fe3—S2^{iv}	92.77 (3)	C3B—C4B—N2B	111.5 (14)
S2^{iii}—Fe3—S2^{iv}	87.23 (3)	C3B—C4B—H17	109.3
S2—Fe3—S2^{iv}	87.23 (3)	N2B—C4B—H17	109.3
S2^{viii}—Fe3—S2^{vi}	87.23 (3)	C3B—C4B—H18	109.3
S2^{iii}—Fe3—S2^{vi}	92.77 (3)	N2B—C4B—H18	109.3
S2—Fe3—S2^{vi}	180.0	H17—C4B—H18	108.0
S2^{iv}—Fe3—S2^{vi}	92.77 (3)	O3—Cl1—O4	110.49 (15)
S2^{viii}—Fe3—S2^{vii}	87.23 (3)	O3—Cl1—O1	109.78 (14)
S2^{iii}—Fe3—S2^{vii}	92.77 (3)	O4—Cl1—O1	109.85 (13)
S2—Fe3—S2^{vii}	92.77 (3)	O3—Cl1—O2	109.71 (13)
S2^{iv}—Fe3—S2^{vii}	179.999 (1)	O4—Cl1—O2	108.82 (14)
S2^{vi}—Fe3—S2^{vii}	87.23 (3)	O1—Cl1—O2	108.16 (12)
N1—Fe2—S1—C2	21.81 (10)	S2^{viii}—Fe3—S2—Fe4	−136.322 (12)
N1^{iv}—Fe2—S1—C2	114.79 (9)	S2^{iii}—Fe3—S2—Fe4	43.680 (12)
S1^{iii}—Fe2—S1—C2	−69.30 (8)	S2^{iv}—Fe3—S2—Fe4	−43.680 (12)
S1^{iv}—Fe2—S1—C2	−157.98 (8)	S2^{vii}—Fe3—S2—Fe4	136.320 (12)
N1—Fe2—S1—Fe1	135.45 (5)	N2^{iii}—Fe4—N2—C4	−175.7 (3)
N1^{iv}—Fe2—S1—Fe1	−131.57 (5)	N2^{iv}—Fe4—N2—C4	91.9 (4)
S1^{iii}—Fe2—S1—Fe1	44.341 (11)	S2^{iii}—Fe4—N2—C4	−88.6 (3)
S1^{iv}—Fe2—S1—Fe1	−44.341 (11)	S2—Fe4—N2—C4	0.1 (3)

S1^{i}—Fe1—S1—C2	−134.32 (9)	Fe4—S2—C3—C4	44.5 (3)
S1^{iii}—Fe1—S1—C2	45.68 (9)	Fe3—S2—C3—C4	119.0 (3)
S1^{iv}—Fe1—S1—C2	132.87 (8)	Fe4—N2—C4—C3	30.6 (5)
S1^{v}—Fe1—S1—C2	−47.13 (8)	S2—C3—C4—N2	−51.4 (5)
S1^{i}—Fe1—S1—Fe2	136.408 (8)	N2B—Fe4—S2B—C3B	27.4 (7)
S1^{iii}—Fe1—S1—Fe2	−43.592 (8)	N2B^{iii}—Fe4—S2B—C3B	119.7 (7)
S1^{iv}—Fe1—S1—Fe2	43.592 (8)	S2B^{iii}—Fe4—S2B—C3B	−153.1 (6)
S1^{v}—Fe1—S1—Fe2	−136.406 (8)	S2B^{iv}—Fe4—S2B—C3B	−64.1 (6)
N1^{iii}—Fe2—N1—C1	−179.13 (15)	N2B—Fe4—S2B—Fe3	136.0 (4)
N1^{iv}—Fe2—N1—C1	−85.94 (19)	N2B^{iii}—Fe4—S2B—Fe3	−131.7 (3)
S1^{iii}—Fe2—N1—C1	93.56 (15)	S2B^{iii}—Fe4—S2B—Fe3	−44.47 (4)
S1—Fe2—N1—C1	4.96 (15)	S2B^{iv}—Fe4—S2B—Fe3	44.47 (4)
Fe2—N1—C1—C2	−37.5 (2)	S2B^{iv}—Fe3—S2B—C3B	42.2 (5)
N1—C1—C2—S1	57.4 (2)	S2B^{iii}—Fe3—S2B—C3B	128.8 (5)
Fe2—S1—C2—C1	−46.89 (16)	S2B^{vii}—Fe3—S2B—C3B	−137.8 (5)
Fe1—S1—C2—C1	−121.55 (14)	S2B^{viii}—Fe3—S2B—C3B	−51.2 (5)
N2—Fe4—S2—C3	−23.37 (18)	S2B^{iv}—Fe3—S2B—Fe4	−43.29 (4)
N2^{iv}—Fe4—S2—C3	−115.61 (18)	S2B^{iii}—Fe3—S2B—Fe4	43.29 (4)
S2^{iii}—Fe4—S2—C3	68.61 (15)	S2B^{vii}—Fe3—S2B—Fe4	136.71 (4)
S2^{iv}—Fe4—S2—C3	157.50 (15)	S2B^{viii}—Fe3—S2B—Fe4	−136.71 (4)
N2—Fe4—S2—Fe3	−136.42 (10)	N2B^{iii}—Fe4—N2B—C4B	−96.9 (14)
N2^{iv}—Fe4—S2—Fe3	131.34 (11)	N2B^{iv}—Fe4—N2B—C4B	170.6 (12)
S2^{iii}—Fe4—S2—Fe3	−44.441 (14)	S2B—Fe4—N2B—C4B	−5.5 (11)
S2^{iv}—Fe4—S2—Fe3	44.441 (14)	S2B^{iv}—Fe4—N2B—C4B	83.3 (11)
S2^{viii}—Fe3—S2—C3	133.80 (16)	Fe4—S2B—C3B—C4B	−54.2 (13)
S2^{iii}—Fe3—S2—C3	−46.20 (16)	Fe3—S2B—C3B—C4B	−127.5 (12)
S2^{iv}—Fe3—S2—C3	−133.55 (16)	S2B—C3B—C4B—N2B	59.4 (18)
S2^{vii}—Fe3—S2—C3	46.45 (16)	Fe4—N2B—C4B—C3B	−32.3 (19)

Symmetry codes: (i) *x*−*y*+2/3, *x*+1/3, −*z*+1/3; (ii) −*x*+2/3, −*y*+4/3, −*z*+1/3; (iii) −*x*+*y*, −*x*+1, *z*; (iv) −*y*+1, *x*−*y*+1, *z*; (v) *y*−1/3, −*x*+*y*+1/3, −*z*+1/3; (vi) −*x*+2/3, −*y*+4/3, −*z*+4/3; (vii) *y*−1/3, −*x*+*y*+1/3, −*z*+4/3; (viii) *x*−*y*+2/3, *x*+1/3, −*z*+4/3.

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N1—H1···O2	0.92	2.28	3.130 (3)	154
N1—H2···O3^{ix}	0.92	2.40	3.162 (3)	140
N2—H7···O1^{iii}	0.92	2.30	3.110 (4)	147
N2—H8···O2	0.92	2.39	3.274 (4)	161
N2B—H13···O2	0.92	2.41	2.984 (12)	121
N2B—H14···O1	0.92	2.27	3.112 (11)	152

Symmetry codes: (ix) *x*−*y*+1/3, *x*+2/3, −*z*+2/3; (iii) −*x*+*y*, −*x*+1, *z*.

Fig. 1

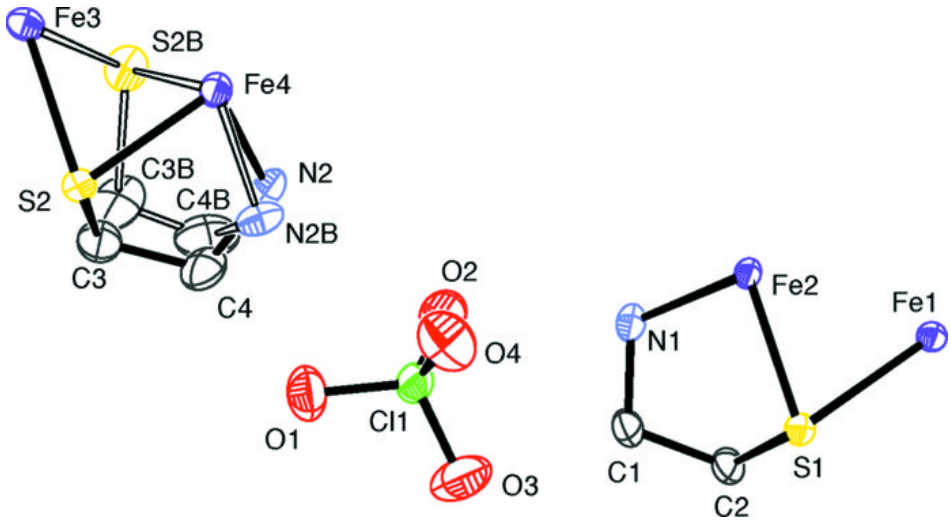
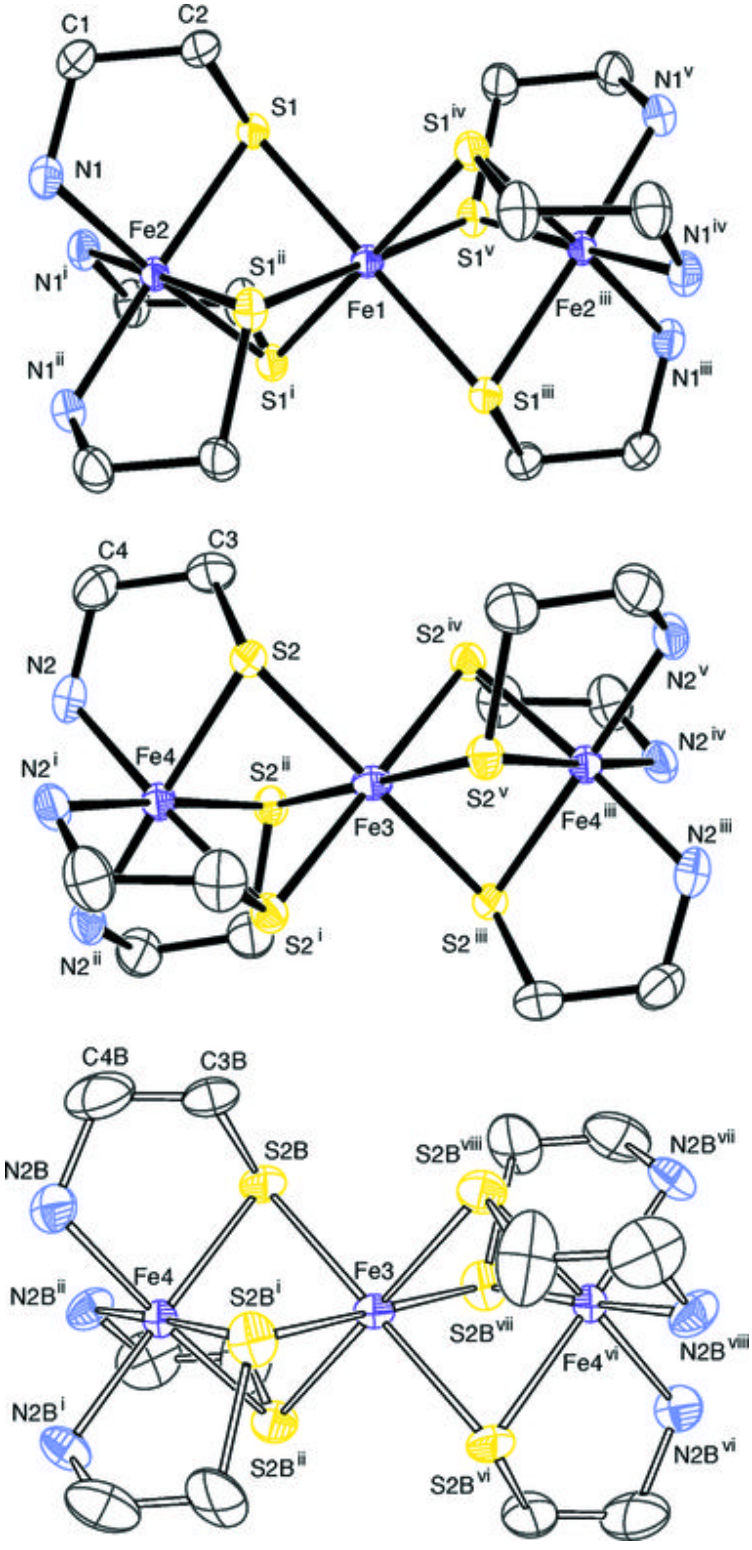

Fe3
S2B
Fe4
C3B
N2
S2
C4B
N2B
C3
C4
O2
O4
Cl1
O1
O3
Fe2
Fe1
N1
C1
S1
C2

Fig. 2


C1
C2
S1
N1
Fe2
N1i
S1ii
N1ii
S1i
Fe1
S1iv
N1v
S1v
N1iv
Fe2iii
N1iii
S1iii
C4
C3
N2
S2
N2i
Fe4
S2ii
N2ii
S2i
Fe3
S2iv
N2v
S2v
N2iv
Fe4iii
N2iii
S2iii
C4B
C3B
N2B
S2B
N2Bii
Fe4
S2Bi
N2Bi
S2Bii
Fe3
S2Bviii
N2Bvii
S2Bvii
N2Bviii
Fe4vi
N2Bvi
S2Bvi